L Number	Hits	Search Text	DB	Time stamp
1	78	((copper cu mn manganese) near2 (Sb antimony)) and (phase adj diagram)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 08:44
-	12	Cu2Sb	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/09 15:01
-	6	Cu".sub.2"Sb	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/09 15:02
-	8	(Cu".sub."2) near Sb	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/09 15:03
-	26	Cu2Sb or Cu".sub.2"Sb or ((Cu".sub."2) near Sb)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/09 15:03
-	7	(Cu2Sb or Cu".sub.2"Sb or ((Cu".sub."2) near Sb)) and battery	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/09 16:07
-	13807	(copper cu mn manganese) near2 (Sb antimony)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/09 16:08
-	2559	((copper cu mn manganese) near2 (Sb antimony)) same (alloy alloyed alloying)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/09 16:08
-	26	(((copper cu mn manganese) near2 (Sb antimony)) same (alloy alloyed alloying)) and (phase adj diagram)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 08:43

L Number	Hits	Search Text	DB	Time stamp
1	33	Cu2As Cu2P	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:10
2	196	copper near2 phosphide	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:10
3	204	copper near2 arsenide	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:11
4	389	(copper near2 arsenide) or (copper near2 phosphide)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:11
5	76	((copper near2 arsenide) or (copper near2 phosphide)) and electrode	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:16
6	26844	(phosphide arsenide phosphorous arsenic) and (lithium li)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:16
7	4682	((phosphide arsenide phosphorous arsenic) and (lithium li)) and electrode	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:17
8	5294	((phosphide arsenide phosphorous arsenic) and (lithium li)) and electrode	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:17
9	381	(((phosphide arsenide phosphorous arsenic) and (lithium li)) and electrode) and nonaqueous	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:19
10	141	((((phosphide arsenide phosphorous arsenic) and (lithium li)) and electrode) and nonaqueous) and (INTERCALAT$4)	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:43
11	49	Fe".sub."2 near As	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:48
12	64	Cr".sub."2 near As	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:55
13	2	LiFeAs	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:56

14	1	LiFeP	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:56
15	2	LiCoAs	USPAT; US-PGPUB; EPO; JPO; DERWENT	2004/08/10 12:57

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YOU HAVE REQUESTED DATA FROM 4 ANSWERS - CONTINUE? Y/(N):y

L8 ANSWER 1 OF 4 CAPLUS COPYRIGHT 2004 ACS on STN
ACCESSION NUMBER: 2000:384637 CAPLUS
DOCUMENT NUMBER: 133:7095
TITLE: Non-aqueous electrolyte secondary **battery** with improved anode
INVENTOR(S): Kasamatsu, Shinji; Yoshizawa, Hiroshi; Okamura, Kazuhiro; Koshina, Hizuru; Shimamura, Harunari; Nitta, Yoshiaki
PATENT ASSIGNEE(S): Matsushita Electric Industrial Co., Ltd., Japan
SOURCE: PCT Int. Appl., 71 pp.
CODEN: PIXXD2
DOCUMENT TYPE: Patent
LANGUAGE: Japanese
FAMILY ACC. NUM. COUNT: 1
PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
WO 2000033401	A1	20000608	WO 1999-JP6687	19991130
W: US				
RW: AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE				
JP 2000173583	A2	20000623	JP 1998-342881	19981202
JP 2000173584	A2	20000623	JP 1998-342882	19981202
JP 2000173585	A2	20000623	JP 1998-342883	19981202
JP 2000173586	A2	20000623	JP 1998-342884	19981202
JP 2000173659	A2	20000623	JP 1998-342895	19981202
JP 2000173660	A2	20000623	JP 1998-342896	19981202
JP 2000173661	A2	20000623	JP 1998-342897	19981202
JP 2000173662	A2	20000623	JP 1998-342898	19981202
JP 2000173615	A2	20000623	JP 1998-342899	19981202
JP 2000173616	A2	20000623	JP 1998-342900	19981202
JP 2000173609	A2	20000623	JP 1998-342901	19981202
JP 2000173610	A2	20000623	JP 1998-342902	19981202
JP 2000173593	A2	20000623	JP 1998-342905	19981202
JP 2000173594	A2	20000623	JP 1998-342906	19981202
EP 1052712	A1	20001115	EP 1999-973175	19991130
R: AT, BE, CH, DE, DK, ES, FR, GB, GR, IT, LI, LU, NL, SE, MC, PT, IE, FI				
US 6605386	B1	20030812	US 2000-601224	20000918
PRIORITY APPLN. INFO.:			JP 1998-342881 A	19981202
			JP 1998-342882 A	19981202
			JP 1998-342883 A	19981202
			JP 1998-342884 A	19981202
			JP 1998-342895 A	19981202
			JP 1998-342896 A	19981202

JP 1998-342897 A 19981202
JP 1998-342898 A 19981202
JP 1998-342899 A 19981202
JP 1998-342900 A 19981202
JP 1998-342901 A 19981202
JP 1998-342902 A 19981202
JP 1998-342905 A 19981202
JP 1998-342906 A 19981202
WO 1999-JP6687 W 19991130

AB The nonaq. electrolyte secondary **battery** comprising pos. and neg. plates capable of absorbing/releasing Li; a nonaq. electrolyte; and a separator; or solid electrolyte, is characterized in that the neg. plate is made chiefly of composite particles of which at least part of the core particles containing ≥1 kind among Sn, Si, and Zn as a constituent element are coated with a solid solution or an intermetallic compound consisting of the constituent element constituting the core particles and a specific element except the constituent element and in that the porosity of the neg. plate mixture layer lies between 10% and 50%. Internal short-circuit between the pos. and neg. plates due to expansion of the neg. plate material can be prevented, the capacity is high, the charging/discharging characteristics are excellent, and rapid charging is possible by adopting the above structure or a structure in which further the amount of nonaq. electrolyte and the thickness of the separator are specified.

IT **Battery** anodes
Battery electrolytes
Secondary **batteries**
Secondary **battery** separators
Short circuits (electrical apparatus)
Solid electrolytes
Solid solutions
Solid state secondary **batteries**
(non-aqueous electrolyte secondary **battery** with improved anode)

IT Intermetallic compounds
RL: DEV (Device component use); NUU (Other use, unclassified); USES (Uses)
(non-aqueous electrolyte secondary **battery** with improved anode)

IT 1302-09-6, Silver selenide (Ag2Se) 1304-76-3, Bismuth oxide (Bi2O3), uses 1304-82-1, Bismuth telluride (Bi2Te3) 1307-96-6, Cobaltous oxide, uses 1308-04-9, Cobalt oxide (Co2O3) 1308-06-1, Cobalt oxide (Co3O4) 1308-38-9, Chromium oxide (Cr2O3), uses 1309-64-4, Antimony oxide (Sb2O3), uses 1313-13-9, Manganese oxide (MnO2), uses 1313-99-1, Nickel oxide (NiO), uses 1314-06-3, Nickel oxide (Ni2O3) 1314-87-0, Lead sulfide (PbS) 1314-91-6, Lead telluride (PbTe) 1315-04-4, Antimony sulfide (Sb2S5) 1315-05-5, Antimony selenide (Sb2Se3) 1317-33-5, Molybdenum disulfide, uses 1317-36-8, Lead oxide (PbO), uses 1317-37-9, Iron sulfide (FeS) 1317-38-0, Cupric oxide, uses 1317-39-1, Cuprous oxide, uses 1317-40-4, Copper sulfide (CuS) 1317-41-5, Copper selenide (CuSe) 1317-42-6, Cobalt sulfide (CoS) 1327-50-0, Antimony telluride (Sb2Te3) 1332-71-4, Cobalt sulfide (Co2S3) 1345-04-6, Antimony sulfide (Sb2S3) 1345-07-9, Bismuth sulfide (Bi2S3) 7439-93-2,

Lithium, uses 7440-21-3, Silicon, uses 7440-31-5, Tin, uses 7440-66-6, Zinc, uses 12002-99-2, Silver telluride (Ag2Te) 12015-75-7, Cobalt sulfide (Co3S4) 12018-06-3, Chromium sulfide (CrS) 12018-22-3, Chromium sulfide (Cr2S3) 12019-23-7, Copper telluride (CuTe) 12019-52-2, Copper telluride (Cu2Te) 12033-33-9, Molybdenum sulfide (Mo2S3) 12053-36-0, Chromium selenide (Cr2Se3) 12059-17-5, Nickel telluride (Ni2Te3) 12063-27-3, Iron sulfide (Fe2S3) 12068-69-8, Bismuth selenide (Bi2Se3) 12068-85-8, Iron sulfide (FeS2) 12069-00-0, Lead selenide (PbSe) 12125-23-4, Manganese sulfide (MnS2) 12137-08-5, Nickel sulfide (Ni2S) 12137-12-1, Nickel sulfide (Ni3S4) 12142-88-0, Nickel telluride (NiTe) 12186-44-6, Iron oxide (FeO4) 12359-48-7, Antimony sulfide (Sb2S4) 12526-66-8, Cobalt selenide (Co3Se4) 16812-54-7, Nickel sulfide (NiS) 18820-29-6, Manganese sulfide (MnS) 20405-64-5, Copper selenide (Cu2Se) 20667-12-3, Silver oxide (Ag2O) 21548-73-2, Silver sulfide (Ag2S) 22205-45-4, Copper sulfide (Cu2S) 39380-32-0, Manganese sulfide (Mn3S4) 51311-17-2, Carbon fluoride 65589-65-3, Cobalt selenide (Co2Se3) 161959-94-0, Antimony selenide (Sb2Se5)
RL: DEV (Device component use); NUU (Other use, unclassified); USES (Uses)
(non-aqueous electrolyte secondary **battery** with improved anode)

REFERENCE COUNT: 18 THERE ARE 18 CITED REFERENCES AVAILABLE FOR THIS RECORD. ALL CITATIONS AVAILABLE IN THE RE FORMAT

L8 ANSWER 2 OF 4 CAPLUS COPYRIGHT 2004 ACS on STN
ACCESSION NUMBER: 2000:313633 CAPLUS
DOCUMENT NUMBER: 132:296154
TITLE: Secondary nonaqueous electrolyte **batteries**
INVENTOR(S): Watanabe, Shoichiro; Fujiwara, Takafumi; Kobayashi, Shigeo
PATENT ASSIGNEE(S): Matsushita Electric Industrial Co., Ltd., Japan
SOURCE: Jpn. Kokai Tokkyo Koho, 9 pp.
CODEN: JKXXAF
DOCUMENT TYPE: Patent
LANGUAGE: Japanese
FAMILY ACC. NUM. COUNT: 1
PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 2000133247	A2	20000512	JP 1998-300550	19981022
PRIORITY APPLN. INFO.:			JP 1998-300550	19981022

AB The **batteries** have a Li intercalating cathode, an anode, and an electrolyte impregnated separator or a solid electrolyte between the electrodes; where the anode active mass contains a compound that electrochem. reduced to form a metal when charging the **battery**. The anode is preferably a Li intercalating carbonaceous material and the compound is a metal sulfide, selenide, and/or telluride.

IT **Battery** anodes
(carbonaceous anodes containing metal chalcogenides for secondary lithium **batteries**)

IT Carbonaceous materials (technological products)
RL: DEV (Device component use); USES (Uses)
(carbonaceous anodes containing metal chalcogenides for secondary lithium **batteries**)
IT 1302-09-6, Silver selenide (Ag2Se) 1304-82-1, Bismuth telluride (Bi2Te3) 1314-87-0, Lead sulfide (PbS) 1314-91-6, Lead telluride (PbTe) 1315-04-4, Antimony sulfide (Sb2S5) 1315-05-5, Antimony selenide (Sb2Se3) 1317-33-5, Molybdenum sulfide (MoS2), uses 1317-37-9, Iron sulfide (FeS) 1317-40-4, Cupric sulfide 1317-41-5, Cupric selenide 1317-42-6, Cobalt sulfide (CoS) 1327-50-0, Antimony telluride (Sb2Te 3) 1332-71-4, Cobalt sulfide (Co2S3) 1345-04-6, Antimony sulfide (Sb2S3) 1345-07-9, Bismuth sulfide (Bi2S3) 11089-54-6, Nickel telluride (Ni2Te) 12002-99-2, Silver telluride (Ag2Te) 12015-75-7, Cobalt sulfide (Co3S4) 12018-06-3, Chromium sulfide (CrS) 12018-22-3, Chromium sulfide (Cr2S3) 12019-23-7, Copper telluride (CuTe) **12019-52-2**, Copper telluride (Cu2Te) 12053-36-0, Chromium selenide (Cr2Se3) 12063-27-3, Iron sulfide (Fe2S3) 12068-69-8, Bismuth selenide (Bi2Se3) 12068-85-8, Iron sulfide (FeS2) 12069-00-0, Lead selenide (PbSe) 12125-23-4, Manganese sulfide (MnS2) 12137-08-5, Nickel sulfide (Ni2S) 12137-12-1, Nickel sulfide (Ni3S4) 12142-88-0, Nickel telluride (NiTe) 12359-48-7, Antimony sulfide (Sb2S4) 12526-66-8, Cobalt selenide (Co3Se4) 16812-54-7, Nickel sulfide (NiS) 18820-29-6, Manganese sulfide (MnS) 20405-64-5, Cuprous selenide 21548-73-2, Silver sulfide (Ag2S) 22205-45-4, Cuprous sulfide 39380-32-0, Manganese sulfide (Mn3S4) 65589-65-3, Cobalt selenide (Co2Se3) 161959-94-0, Antimony selenide (Sb2Se5)
RL: MOA (Modifier or additive use); USES (Uses)
(carbonaceous anodes containing metal chalcogenides for secondary lithium **batteries**)

L8 ANSWER 3 OF 4 CAPLUS COPYRIGHT 2004 ACS on STN
ACCESSION NUMBER: 1995:410542 CAPLUS
DOCUMENT NUMBER: 122:165510
TITLE: Secondary solid state lithium **batteries**
INVENTOR(S): Takada, Kazunori; Iwamoto, Kazuya; Aotani, Noboru; Kondo, Shigeo
PATENT ASSIGNEE(S): Matsushita Electric Ind Co Ltd, Japan
SOURCE: Jpn. Kokai Tokkyo Koho, 7 pp.
CODEN: JKXXAF
DOCUMENT TYPE: Patent
LANGUAGE: Japanese
FAMILY ACC. NUM. COUNT: 1
PATENT INFORMATION:

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
JP 06275312	A2	19940930	JP 1993-87985	19930322
PRIORITY APPLN. INFO.:			JP 1993-87985	19930322

AB The **batteries** use Li+ conductive solid electrolytes and Cu chalcogenide cathodes. The chalcogenide may be Cu2S, Cu2Se, Cu2Te and

their derivs. and the cathodes may also contain Li chalcogenides and the electrolyte. These **batteries** ae suitable for high rate charge and discharge.

IT **Batteries**, secondary
(secondary lithium **batteries** with copper chalcogenide cathodes and lithium ion conductive solid electrolytes)

IT **Battery** electrolytes
(secondary lithium **batteries** with lithium ion conductive solid electrolytes)

IT Cathodes
(**battery**, secondary solid state lithium **batteries** with cathodes containing copper and chalcogens)

IT 161286-52-8, Lithium sulfide thiosilicate (Li1.2S0.2(SiS3)0.4) 161286-53-9, Lithium phosphenotrithioate sulfide (Li1.2(PS3)0.8S0.2) 161286-54-0, Lithium phosphate sulfide thiosilicate (Li1.24(PO4)0.02S0.2(SiS3)0.39) 161487-41-8, Lithium iodide thiosilicate (LiI0.3(SiS3)0.35)
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(secondary lithium **batteries** with lithium ion conductive solid electrolytes)

IT 7440-50-8, Copper, uses 7704-34-9, Sulfur, uses 7782-49-2, Selenium, uses
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(secondary solid state lithium **batteries** with cathodes containing copper and chalcogens)

IT **12019-52-2**, Cuprous telluride 12345-69-6, Chromium copper sulfide (CrCuS2) 20405-64-5, Cuprous selenide 22205-45-4, Cuprous sulfide
RL: DEV (Device component use); TEM (Technical or engineered material use); USES (Uses)
(secondary solid state lithium **batteries** with copper chalcogenide cathodes)

L8 ANSWER 4 OF 4 CAPLUS COPYRIGHT 2004 ACS on STN
ACCESSION NUMBER: 1977:147647 CAPLUS
DOCUMENT NUMBER: 86:147647
TITLE: Solid-state ionics-solid electrolyte cells with copper ion conductors
AUTHOR(S): Takahashi, T.; Yamamoto, O.
CORPORATE SOURCE: Fac. Eng., Nagoya Univ., Nagoya, Japan
SOURCE: Journal of Applied Electrochemistry (1977), 7(1), 37-43
CODEN: JAELBJ; ISSN: 0021-891X
DOCUMENT TYPE: Journal
LANGUAGE: English

AB A solid electrolyte cell was developed using a high Cu ion conducting solid electrolyte, 7CuBr·C6H12N4CH3Br (N-methylhexamethylenetetraminium bromide), a Cu anode, and a chalcogen

cathode. The open-circuit voltages of the cells with S, Se, and Te as cathodes were 0.448, 0.373, and 0.258 V, resp. at 25°. These cells yielded a current of several tens of μA at room temperature and several mA at 114° without appreciable polarization. An energy d. of 4.5 Wh kg-1 at room temperature was evaluated from the wts. of the electrolyte and electrode materials for the cell using a Se-cathode in the discharge c.d. range 60-150 μA cm-2.

IT Group VIA elements
RL: USES (Uses)
(cathodes, in cell with copper anode and copper methylhexamethylenetetraminium bromide electrolyte)

IT **Batteries**, primary
(copper-chalcogenide, with copper-methylhexamethylenetetraminium bromide electrolyte)

IT Heat of formation
(of copper chalcogenides)

IT Electric conductivity and conduction
(of copper methylhexamethylenetetraminium bromide electrolyte)

IT Electric potential
(of copper-chalcogen cell with copper methylhexamethylenetetraminium bromide electrolyte, thermodn. in relation to)

IT Electrolytic polarization
(of copper-chalcogenide cell with copper methylhexamethylenetetraminium bromide electrolyte)

IT Entropy
(of formation, of copper chalcogenides)

IT 7440-50-8, uses and miscellaneous
RL: USES (Uses)
(anode, in cell with chalcogen cathode and copper-methylhexamethylenetetraminium bromide electrolyte)

IT 7704-34-9, uses and miscellaneous 7782-49-2, uses and miscellaneous 13494-80-9, uses and miscellaneous
RL: USES (Uses)
(cathodes, in cell with copper anode and copper methylhexamethylenetetraminium bromide electrolyte)

IT 57384-10-8
RL: PRP (Properties)
(electrolyte, for cells with chalcogen cathodes and copper anode)

IT **12019-52-2** 20405-64-5 22205-45-4
RL: PRP (Properties)
(thermodn. of, elec. potential in relation to)

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     (FILE 'HOME' ENTERED AT 12:59:09 ON 10 AUG 2004)

     FILE 'REGISTRY' ENTERED AT 12:59:14 ON 10 AUG 2004
                  E ASFE2/MF
L1              1 S E3
                  E ASCR2/MF
L2              1 S E3
                  E CU4TE2/MF
                  E CU2TE/MF
L3              2 S E3
                  E ASFELI/MF
                  E ASFE/MF
L4              4 S E3-4
                  E FEP/MF
L5              1 S E5
                  E ASCO/MF
L6              1 S E4

     FILE 'CAPLUS' ENTERED AT 13:04:12 ON 10 AUG 2004
L7            647 S L1-L6
L8              4 S L7 AND BATTER###
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13 Articles Found

pub-date > 1979 and Cu2Sb AND structure AND battery

1. ☐ **Electrochemical lithiation and delithiation of $FeSb_2$ anodes for lithium-ion batteries** • SHORT COMMUNICATION
Materials Letters, Volume 57, Issue 30, December 2003, Pages 4673-4677
J. Xie, X. B. Zhao, G. S. Cao, M. J. Zhao, Y. D. Zhong and L. Z. Deng
Abstract

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2. ☐ **Alternative anode materials for lithium-ion batteries: a study of Ag_3Sb** • ARTICLE
Journal of Power Sources, Volumes 119-121, 1 June 2003, Pages 64-68
J. T. Vaughey, L. Fransson, H. A. Swinger, K. Edström and M. M. Thackeray
SummaryPlus | Full Text + Links | PDF (168 K)

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3. ☐ **Inorganic materials for the negative electrode of lithium-ion batteries: state-of-the-art and future prospects** • REVIEW ARTICLE
Materials Science and Engineering: R: Reports, Volume 40, Issue 3, 14 February 2003, Pages 103-136
José L. Tirado
Abstract

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4. ☐ **Ex-situ XRD studies of $CoSb_3$ compound as the anode material for lithium ion batteries** • ARTICLE
Journal of Electroanalytical Chemistry, Volume 542, 30 January 2003, Pages 1-6
Jian Xie, Xinbing Zhao, Gaoshao Cao, Yaodong Zhong and Mingjian Zhao
SummaryPlus | Full Text + Links | PDF (315 K)

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5. ☐ **Structural considerations of intermetallic electrodes for lithium batteries** • ARTICLE
Journal of Power Sources, Volume 113, Issue 1, 1 January 2003, Pages 124-130
M. M. Thackeray, J. T. Vaughey, C. S. Johnson, A. J. Kropf, R. Benedek, L. M. L. Fransson and K. Edstrom

SummaryPlus | Full Text + Links | PDF (402 K)

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6. ☐ **Lithium reactions with intermetallic-compound electrodes** • ARTICLE
Journal of Power Sources, Volume 110, Issue 2, 22 August 2002, Pages 406-411
R. Benedek and M. M. Thackeray
SummaryPlus | Full Text + Links | PDF (122 K)

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7. ☐ **Recent advances in the US Department of Energy's energy storage technology research and development programs for hybrid electric and electric vehicles** • ARTICLE
Journal of Power Sources, Volume 110, Issue 2, 22 August 2002, Pages 471-474
Irwin B. Weinstock
SummaryPlus | Full Text + Links | PDF (162 K)

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8. ☐ **Phase transitions in lithiated Cu_2Sb anodes for lithium batteries: an in situ X-ray diffraction study** • SHORT COMMUNICATION
Electrochemistry Communications, Volume 3, Issue 7, July 2001, Pages 317-323
L. M. L. Fransson, J. T. Vaughey, R. Benedek, K. Edström, J. O. Thomas and M. M. Thackeray
Abstract

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9. ☐ **Lead-free Solders in Microelectronics** • ARTICLE
Materials Science and Engineering: R: Reports, Volume 27, Issues 5-6, 1 June 2000, Pages 95-141
Mulugeta Abtew and Guna Selvaduray
Abstract

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10. ☐ **Synthesis and characterization of a new copper iron telluride** • ARTICLE
Journal of Alloys and Compounds, Volume 217, Issue 2, 1 February 1995, Pages 250-252
Jaime Llanos and Carlos Mujica
Abstract | Abstract + References | PDF (218 K)

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11. ☐ **Structure and chemical bonding in zintl-phases containing lithium** • REVIEW ARTICLE
Progress in Solid State Chemistry, Volume 20, Issue 1, 1990, Pages 1-45
Reinhard Nesper
Abstract

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12. ☐ **6. Titanium** • ARTICLE
Coordination Chemistry Reviews, Volume 73, October 1986, Pages 175-279
Mary A. Jamieson, Nick Serpone[1] and Ezio Pelizzetti
Abstract

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13. ☐ **Sensing devices for sodium vapour detection** • ARTICLE
Sensors and Actuators, Volume 9, Issue 4, July 1986, Pages 323-331

Patrick Kayser, Jean Claude Launay, Bernard Tanguy, Francis Menil and Josik PortierClaude Lucat
Abstract

13 Articles Found

pub-date > 1979 and Cu2Sb AND structure AND battery

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